- **96.(A)** A silica garden or chemical garden is an experiment normally performed by adding coloured metals salt such as CuSO₄ or CoCl₂ to an aqueous solution of sodium silicate. This results in growth of plant like forms.
- **97.(A)** Silicon carbide is a hard substance hence used as abrasive.
- **98.(D)** Because some of the silicon chloride gets hydrolyzed.
- **99.(A)** Silicon have large number of alkyl groups which have appreciable hydrophobic properties. As general representation of linear silicon is

$$\begin{array}{c|c} R & R & R \\ | & R & Si - O - Si - O - Si - O - H \\ | & R & R & R \end{array}$$

$$\begin{array}{c|c} R & R & R & R \\ | & R & R & R \end{array}$$

$$\begin{array}{c|c} R & R & R & R & R \\ | & R & R & R & R \end{array}$$

$$\begin{array}{c|c} R & R & R & R & R & R & R \\ | & R & R & R & R & R & R \end{array}$$

 $\begin{array}{c|c} R & R & R \\ | & Si-O \\ | & Si-O \\ | & R \end{array} \begin{array}{c} R & R \\ | & | \\ Si-O-Si-R \\ | & | \\ R & R \end{array}$

- **100.(BC)** O-C-O bond angle is 180° due to sp hybridization of carbon and also due to p-p bonding between carbon and oxygen.
- 101.(AB) Organic silicone chain shows condensation reaction with R₃SiCl as shown

- 102.(BCD) Graphite is thermodynamically more stable than Diamond because of
 - (1) Double bond character in C C bond
 - (2) e⁻ are delocalized in a manner that aromaticity get introduced in graphite sheets.
 - (3) Graphite has greater entropy which is because of
 - (a) Layers are much distant
 - (b) Layers can get align randomly in AB AB packing and ABC ABC packing [XIIth Class Solid State]

C₆₀, Buckminster has cage like structure.

Due to slippery nature of graphite. It is used as dry lubricant : Other such example is molybdenum disulfide.

103.(BD)
$$O = C = O \longleftrightarrow O = C - O$$
104.(C) $Si + HF \longleftrightarrow H_2SiF_6$

105.(B) Gas P and Q are same i.e. H₂